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CF₄
(carbon tetrafluoride)

A. The Lewis Structure of CF₄

1) Valence electrons	32	4	7
2) Required electrons	40	8	8
3) Bonding electrons	8		
4) Non-bonding electrons	24		

C is less electronegative than F. So, C becomes the central atom in our draft. Assign single bonds between the central atom and terminal atoms.
[It uses \times for the bonds to show the e^- used. This will be changed to $—$ in the final structure.]

Underline each compound below that is nonpolar. Draw a Lewis structure for each and label with the molecular geometry and polarity (polar or nonpolar) to help answer this question.

Carbon dioxide (CO₂), glucose (C₆H₁₂O₆), or molecular oxygen (O₂)

ESTRUCTURA DE LEWIS AcademiaVasquez

HBrO

H : 1s¹
Br : 1s² 2s² 2p⁶ 3s² 3p⁶ 4s² 3d¹⁰ 4p⁵
O : 1s² 2s² 2p⁴

1	s			
2	s	p		
3	s	p	d	
4	s	p	d	f
5	s	p	d	f
6	s	p	d	
7	s	p	d	

Carbon Tetrabromide CBr₄

C → 4e⁻ = 4e⁻
Br → 7e⁻ x 4 = 28e⁻
Total = 32e⁻

6 + 6 + 6 + 6 + 8 = 32e⁻

octet rule ✓

COCl₂

e- wanted 32
valence e- 24
bonding e- 8
8/2 = 4

(A) SF₄ has a trigonal bipyramidal arrangement of the five regions of electronic density. Examples include H₂S and NH₃. It is possible to view or hide the moments of bond, the molecular dipole and the partial charges on the right. The basic geometry is trigonal planar with bond corners at 120°, but we see that the double bond causes slightly larger corners (121°) and the angle between the individual bonds is slightly small (118°). Why is the corner of H-C-H in NH₃ the smallest of the bond corners? H-C-H in CH₄? BCl₃ also has a Trigonal planar molecular structure (Figure 8). (b) A is very electronegative and B and C are not. The BeF₂ molecule adopts a linear structure in which the two bonds are the most distant possible, on the opposite sides of the Be atom. (c) One of the regions is a lone pair, which translates into a swing molecular structure. Two regions of electronic density around a central atom in a molecule form linear geometry; Three regions form a trigonal planar geometry; Four regions form a tetrahedral geometry; Five regions form a trigonal bipyramidal geometry; And six regions form an octahedral geometry. Any molecule with five pairs of electrons around the central atoms including no solitary couple will be trigonal bipyramidal. The ideal bond corners in a trigonal pyramid are based on the geometry of the torque of tetrahedral electrons. We can also use the "Name" control boxes at the bottom left to view or hide the geometry of the electron pair (called "electronic geometry" in the simulator) and/or molecular structure (called "molecular geometry" in the simulator). (b) The molecular structure is flat with the lone pairs directly facing each other. The electrostatic repulsion of these electrons is reduced when the various regions of electronic density assume positions as far as possible from each other. Explain the difference observed. Count the number of regions of electronic density around the central atom. In the bond angle is 104.5°. Predict the electron pair geometry and the molecular structure of each of the following molecules or ions: (a) IF₆⁺ (b) CF₄ (c) BF₃ (d) SiF₅⁻ (e) BeCl₂ What are the electron-pair geometry and the molecular structure of each of the following molecules or ions? Chemistry End of Chapter Exercises Explain why the HOH molecule is bent, whereas the HBeH molecule is linear. The order of electron-pair repulsions from greatest to least repulsion is: lone pair-lone pair > lone pair-bonding pair > bonding pair-bonding pair. This order of repulsions determines the amount of space occupied by different regions of electrons. Is X beryllium or sulfur? The two lone pairs are on opposite sides of the octahedron (180° apart), giving a square planar molecular structure that minimizes lone pair-lone pair repulsions (Figure 6). Use the electronegativity controls to determine how the molecular dipole will look for the starting bent molecule H-F. (a) A and C are very electronegative and B is in the middle of the range. (c) Both the electron geometry and the molecular structure are linear. The largest bond moments will occur with the largest partial charges. Although C and S have very similar electronegativity values, S is slightly more electronegative than C, and so the C-S bond is just slightly polar. Build the molecule HCN in the simulator based on the following Lewis structure: [H]-C≡N. Click on each bond type or lone pair at right to add that group to the central atom. According to VSEPR theory, the terminal atom locations (Xs in Figure 6) are equivalent within the linear, trigonal planar, and tetrahedral electron-pair geometries (the first three rows of the table). Only ClF₃, ClO₂, SeF₄, and PH₂ have dipole moments. All of these molecules and ions contain polar bonds. (b) The trigonal pyramidal molecular structure is determined from the electron-pair geometry. The magnitude of a bond dipole moment is represented by the Greek letter μ (mu) and is given by the formula shown here, where Q is the magnitude of the partial charges (determined by the electronegativity difference) and r is the distance between the charges: μ = Qr. This bond moment can be represented as a vector, a quantity having both direction and magnitude (Figure 13). (b) With a partial positive charge on A, turn on the electric field and describe what happens. Chloromethane, CH₃Cl, is another example of a polar molecule. Then try to find a chemical formula that would match the structure you have drawn. (a) Each CO bond has a bond dipole moment, but they point in opposite directions so that the net CO₂ molecule is nonpolar. The Lewis structure of BeF₂ (Figure 2) shows only two electron pairs around the central beryllium atom. (b) Both the electron geometry and the molecular structure are trigonal bipyramidal. For HF, there is a larger dipole moment because there is a larger difference in electronegativity. Bond distances are measured in Ångströms (1 Å = 10⁻¹⁰ m) or picometers (1 pm = 10⁻¹² m, 100 pm = 1 Å). (c) With a small partial negative charge on A, turn on the electric field and describe what happens. Check Your Learning Predict the electron pair geometry and molecular structure for molecules of XeF₂. The two solutions above represent how unevenly the electrons are shared in the bond. You may need to rotate the molecules in three dimensions to see certain dipoles. The electron-pair geometry is trigonal bipyramidal. Is X boron or phosphorus? Draw the Lewis electron dot structures for these molecules, including resonance structures where appropriate: (a) CS₂

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